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Anion-Initiated Trifluoromethylation by TMSCF₃: Deconvolution of the Siliconate-Carbanion Dichotomy by Stopped-Flow NMR/IR

Craig P. Johnston,^{†‡} Thomas H. West,^{†‡} Ruth E. Dooley,[†] Marc Reid,[†] Ariana B. Jones,[†] Edward J. King,[‡] Andrew G. Leach,[‡] and Guy C. Lloyd-Jones^{*,†}

[†] EaStChem, University of Edinburgh, Joseph Black Building, David Brewster Road, Edinburgh, EH9 3FJ, UK

[‡] School of Pharmacy and Biomolecular Sciences, Liverpool John Moores University, Byrom Street, Liverpool, L3 3AF, UK

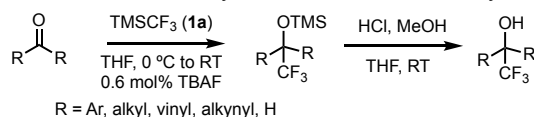
^{*} TgK Scientific Limited, 7 Long's Yard, St Margaret's Street, Bradford-on-Avon, BA15 1DH, UK

ABSTRACT: The mechanism of CF₃ transfer from R₃SiCF₃ (R = Me, Et, *i*Pr) to ketones and aldehydes, initiated by M⁺X⁻ (<0.004 to 10 mol%) has been investigated by analysis of kinetics (variable-ratio stopped-flow NMR and IR), ¹³C/²H KIEs, LFER, addition of ligands (18-c-6, crypt-222), and density functional theory (DFT) calculations. The kinetics, reaction orders, and selectivity vary substantially with reagent (R₃SiCF₃) and initiator (M⁺X⁻). Traces of exogenous inhibitors present in the R₃SiCF₃ reagents, which vary substantially in proportion and identity between batches and suppliers, also affect the kinetics. Some reactions are complete in milliseconds, others take hours, others stall before completion. Despite these differences, a general mechanism has been elucidated in which the product alkoxide and CF₃⁻ anion act as chain carriers in an anionic chain reaction. Silyl enol ether generation competes with 1,2-addition and involves protonation of CF₃⁻ by the α-C–H of the ketone, and the OH of the enol. The overarching mechanism for trifluoromethylation by **1**, in which pentacoordinate siliconate intermediates are unable to directly transfer CF₃⁻ as a nucleophile or base, rationalizes why the turnover rate (per M⁺X⁻ initiator) depends on the initial concentration (but not identity) of X⁻, the identity (but not concentration) of M⁺, the identity of the R₃SiCF₃ reagent, and the carbonyl / R₃SiCF₃ ratio. It also rationalizes which R₃SiCF₃ reagent effects the most rapid trifluoromethylation, for a specific M⁺X⁻ initiator.

INTRODUCTION

The inclusion of fluorine-substituents in organic molecules is of pivotal importance to developments in, inter alia, pharmaceuticals,¹ agrochemicals,² electronics,³ materials chemistry,⁴ polymers,⁵ synthesis,⁶ and catalysis.⁷ The transfer of a formally-nucleophilic CF₃-moiety to an electrophile is a preeminent method for the synthesis of trifluoromethylated compounds.⁸ Conditions range from base-mediated reactions with fluoroform (CF₃H),⁹ through to finely-tuned borazine-based CF₃-carriers recently reported by Szymczak.¹⁰ In 1989, Ruppert reported that TMSCF₃ (**1a**)¹¹ undergoes addition to aldehydes and ketones in the presence of 10 mol% KF.¹² A faster process, using a soluble initiator (Bu₄NF.xH₂O; 0.6 mol%, 'TBAF') was reported soon after, by Prakash and Olah.¹³ Acidic work-up affords the corresponding trifluoromethylated alcohols in good yield, Scheme 1.

Scheme 1. Trifluoromethylation of ketones / aldehydes.^{12,13a}



This mild and selective process¹⁴ swiftly became adopted for the preparation of trifluoromethyl-carbinols,¹⁵ including enantioselective additions involving enantiopure ammonium

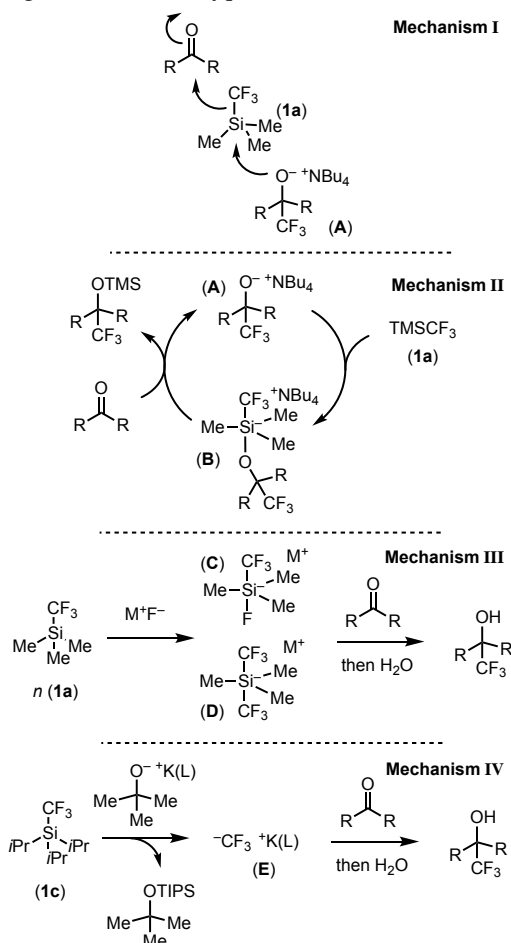
salts as initiators.¹⁶ Indeed, over the last decade there has been an explosion of interest¹⁷ in CF₃ transfer from TMSCF₃ (**1a**) to carbon (e.g. carbonyls,¹⁴⁻¹⁷ imines,¹⁸ vinylhalides¹⁹ and aromatics²⁰) and to heteroatoms such as sulfur,²¹ selenium,²² phosphorus,²³ boron,²⁴ iodine,²⁵ and bismuth.²⁶ The formal loss of fluoride from CF₃ to facilitate electrophilic TMSCF₂-transfer,²⁷ or carbenoid CF₂-transfer,²⁸ has also been developed, as have numerous metal-mediated and catalyzed processes involving CF₃ derived from TMSCF₃ (**1a**).²⁹

Despite anion-initiated trifluoromethylation by **1a** having become a mainstream synthetic method,¹⁷⁻²⁶ surprisingly little detail has emerged on the mechanism of CF₃ transfer, under the conditions of application, Scheme 1.³⁰ Various mechanistic dichotomies, including, inter alia, fluoride-initiation versus fluoride catalysis, and siliconate versus carbanion^{23a} pathways, have been noted by Denmark,^{30a} and by Reich,^{30b} both of whom emphasize the lack of salient kinetic data.

Herein we report the first detailed study of the mechanism of anion-initiated CF₃ transfer from TMSCF₃ (**1a**) to ketone and aldehyde electrophiles.^{12,17} The *in situ* NMR/IR investigations include analysis of reaction kinetics, selectivity, and side reactions, and the contrasting behaviour of homologues TES- (**1b**) and TIPS- (**1c**). Throughout the investigation, the kinetic studies have both informed, and been directed by,

density functional theory (DFT) analysis of proposed intermediates. What emerges is a nuanced kinetic landscape in which trifluoromethyl transfer proceeds via a carbanion pathway (CF_3^-), with the rate dictated by the identity of the electrophile, the concentration of the initiating anion, the identity of the initiator counter-cation, the electrophile / R_3SiCF_3 (**1**) concentration ratio, and the identity of the reagent (**1a-c**).

Scheme 2. Mechanisms (I–IV) for anion-induced trifluoromethylation of ketones using Ruppert's reagent (**1a**)¹¹ and homologues. L = 18-c-6, crypt-222. See text for full discussion.



RESULTS AND DISCUSSION

1. Prior Studies. In early studies, a termolecular anionic chain-reaction (mechanism I, Scheme 2), was suggested for trifluoromethylation by **1a**.^{13a} This was later expanded to a two-step process (mechanism II), where a pentacoordinate alkoxy-siliconate (**B**) delivers CF_3 to the ketone, and in doing so liberates the *O*-silylated product.¹⁴ Mechanism II has been extensively adopted in the design and interpretation of asymmetric trifluoromethylation.^{16,29,31}

In 1999, Naumann,³² and Kolomeitsev and Röschenhalter³³ independently reported on the reaction of a range of soluble fluoride sources (e.g. $[\text{Me}_4\text{N}]^+\text{F}^-$) with TMSCF_3 (**1a**) at low temperature. Detailed ^1H , ^{13}C , ^{19}F and ^{29}Si NMR analysis identified the products as pentacoordinate complexes $[\text{Me}_3\text{Si}(\text{F})(\text{CF}_3)]^-\text{M}^+$ (**C**) and $[\text{Me}_3\text{Si}(\text{CF}_3)_2]^- \text{M}^+$ (**D**). Both complexes decompose above -20°C .^{32,34} The speciation (**C** / **D**) is

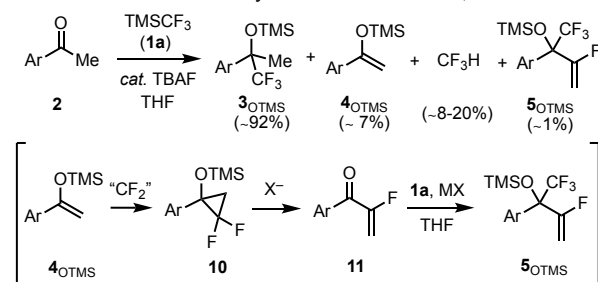
dependent on the stoichiometry (M^+F^- / **1a**), and the structure of **D** was confirmed by single crystal X-ray diffraction. Addition of cyclohexanone at -60°C , followed by hydrolysis, afforded the corresponding trifluoromethylated alcohol, mechanism III.^{32,35}

In 2014, Prakash³⁶ showed that the elusive³⁷ trifluoromethyl anion(oid)³⁸ can be detected *in situ* (^{13}C , ^{19}F NMR) at low temperatures after addition of KO^tBu / 18-crown-6 to **1a**. With the much bulkier reagent TIPSCF_3 (**1c**), the generation of ion-paired $[\text{K}(18\text{-c-}6)]^+[\text{CF}_3]^-$ (**E**) proceeds quantitatively at -78°C over a period of 30 mins. Subsequent addition of PhCOMe (11 equiv.) or PhCHO (4 equiv.) afforded CF_3 -addition products (22–68%) after quenching with H_2O , mechanism IV.³⁶ In 2015, Grushin³⁸ demonstrated that use of crypt-222 (L, Scheme 2) facilitates generation of the free CF_3^- carbanion; a THF solution-phase "noncovalently-bound ionic species".³⁸ The structure of the highly air- and temperature-sensitive salt, $[\text{K}(\text{crypt-222})]^+[\text{CF}_3]^-$ (**E**), was confirmed by single crystal X-ray diffraction.^{38,39}

The pioneering studies summarized above have been highly enlightening regarding the structure and stability of penta-coordinate (trifluoromethyl)siliconates (**C**, **D**),^{32,33} and their ability to release the trifluoromethane anion(oid) (**E**) under specific conditions.^{36,38} However, they do not yield direct detail on the kinetics and mode of transfer of CF_3 from TMSCF_3 (**1a**) to a carbonyl electrophile, using a catalytic fluoride-based initiator (M^+X^-), at ambient temperature.^{12,13}

2. Preliminary Investigations. We began by study of the reaction of TMSCF_3 (**1a**) with aldehydes and ketones in THF, chlorobenzene, and DMF. After addition of catalytic quantities (0.1 to 1 mol%) of TBAF, ^{19}F NMR readily facilitated analysis of the proportions of residual reagent (**1a**) and the [1,2]-addition products. The reaction of 4-fluoroacetophenone (**2**) in THF at ambient temperatures proved ideal, the additional ^{19}F nucleus allowing simultaneous analysis of reagent (**1a**; 0.48 M), substrate (**2**; 0.40 M), and product (**3OTMS**), Scheme 3.

Scheme 3. Trifluoromethylation of ketone **2** ($\text{Ar} = 4\text{-F-C}_6\text{H}_4$).



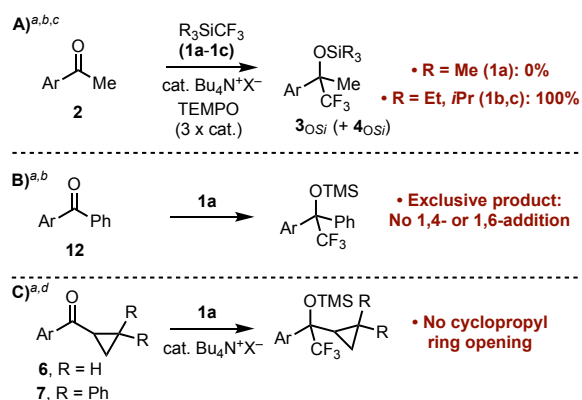
Reactions were assembled manually in 5 mm NMR tubes in the glove box prior to analysis *in situ* by ^{19}F NMR. Three side-products were identified: fluoroform (CF_3H), the silylenol ether (**4OTMS**) and a homologated addition product **5OTMS**. Reactions conducted in *ds*-THF proceeded analogously, and generated CF_3H , not CF_3D .⁴⁰ The identity of **5OTMS**, which was confirmed by independent synthesis, is consistent with difluorocyclopropanation of silylenol ether **4OTMS** to generate **10**, followed by a known⁴¹ anion-induced ring-opening elimination to give fluoroenone **11**, and subsequent 1,2-selective¹²

trifluoromethylation. Addition of independently synthesized⁴² **10** to the reaction (Scheme 3) generated **5_{OTMS}**.

Reaction rates, and extent of fluoroform generation (Scheme 3), were found to vary significantly between batches of TBAF (1 M, THF, ~5 wt% H₂O). Replacing TBAF with anhydrous [Bu₄N][Ph₃SiF₂] ('TBAT')⁴³ gave more reproducible data. However, the fast turnover precluded detailed kinetic analysis; this aspect was addressed using stopped-flow methods, *vide infra*. Nonetheless, ¹⁹F NMR analysis revealed that CF₃H is liberated in two distinct phases. The first is an initial burst of extremely rapid CF₃H generation, and arises from TBAT-catalyzed reaction of TMSCF₃ (**1a**) with traces of adventitious water.⁴⁴ The second phase of CF₃H generation proceeds in concert with reaction of the ketone (**2**) and directly correlates with the rate of generation of silylenol ether (**4_{OTMS}**), as confirmed by ²H-labelling (*d*₃-**2** → CF₃D + *d*₂-**4_{OTMS}**). The selectivity (**3_{OTMS}** versus **4_{OTMS}**) is discussed later.

3. Stability, Inhibition, and Tests for Radicals. The stability of the reaction system after complete consumption of the limiting reagent (ketone **2** or TMSCF₃ **1a**) was found to depend on which one was in excess. Reactions in which **2** was in excess, underwent turn-over on addition of further TMSCF₃ **1a**, even after a period of many hours. In contrast, for reactions where **1a** was in excess, additional **2** had to be added within a few minutes to fully reinstate turnover (see SI); consistent with the known instability of pentacoordinate (trifluoromethyl)siliconates, e.g **C** and **D**, at ambient temperatures.³²⁻³⁴ Further tests established that the reactions were not sensitive to exogenous water *per se*, as they rapidly self-dehydrated via generation of CF₃H + hexamethyldisiloxane, prior to reaction of the ketone (**2**).⁴⁵ The rates were unaffected by visible light, by exogenous product (**3_{OTMS}**), and by CF₃H. Deliberate sparging of the normal reaction mixture (**1a** / **2** / TBAT 0.15 mM, 0.038 mol%, THF, Scheme 3) with air caused complete inhibition of turnover, but only when a sufficient volume of CO₂ (~400 ppm) had been added to convert the active anion(s) into trifluoroacetate (i.e. [Bu₄N][CF₃CO₂], detected by ¹⁹F NMR). Separate controls confirmed that the rate of trifluoromethylation is unaffected by CO₂-scrubbed air, and that [Bu₄N][CF₃CO₂] is not effective as an initiator.

Scheme 4. Tests for radical intermediates.



^a(Ar = 4-F-C₆H₄); ketone (**2**, **6**, **7**, **12**, 0.40 M), **1a-c** (0.48 M), THF, 21 °C. ^bTBAT (0.04 mol%, 0.15 mM), ^cTEMPO (0.12 mol%, 0.45 mM; up to 80 mM with **1b,c**). ^dTBAF (4 mM, 1 mol%).

However, the reactions were inhibited by addition of the persistent radical, TEMPO. Indeed, just 0.45 mM TEMPO induced complete inhibition of the reaction of **1a** with **2**, initiated by 0.15 mM TBAT (Scheme 4, A). In contrast, TEMPO had negligible impact on reactions employing TES (**1b**) and TIPS (**1c**), even when present at much higher concentrations (80 mM TEMPO); the origins of this profound difference in behavior is discussed later. Nonetheless, further tests for discrete radical intermediates^{46,47} were conclusively negative: 4-F-benzophenone (**12**) exclusively underwent 1,2-addition (Scheme 4, B),^{48,49} cyclopropyl ketones (**6/7**) reacted without any trace of competing ring-opening⁵⁰ (Scheme 4, C), and competition between ketone **2** and 4-biphenyl methylketone for limiting TMSCF₃ (**1a**) favored **2** ($k_{rel} = 1.93$).⁵¹

4. General Effects of Initiator on Rate and Selectivity. A range of initiators (M^+X^-) were tested and found to strongly impact the reaction outcome. In the majority of cases, the reactions initiated 'instantly' and the identity of X^- had no influence on the rate⁵² or selectivity (**3_{OTMS}** / **4_{OTMS}**). Specific effects were found to be dictated by the identity of the counter-cation (M^+), Table 1.

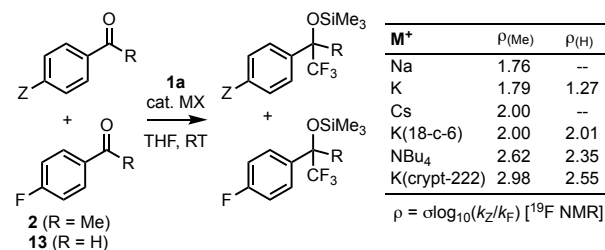
Table 1. Examples of effect of initiator M^+ and reagent (**1a-c**) on selectivity (**3_{OSi}** / **4_{OSi}**) and rate of trifluoromethylation of **2**.

M^+	$[M^+X^-]_0$ mM	TMS 1a 3/4^a (time) ^b	TES 1b 3/4^a (time) ^b	TIPS 1c 3/4^a (time) ^b
[Bu ₄ N] ⁺	1.5	12 / 1 (<90 s)	1.5 / 1 (<90 s)	1 / 1 (30 min.)
[K] ⁺	0.15	36 / 1 (<90 s)	3.0 / 1 ^c (30 min.)	NR
[K(L)] ⁺ ^d	1.5	6.6 / 1 (6 min.)	2.4 / 1 (<90 s)	1 / 1 (3.6 min.)

^aSelectivity **3_{OSi}** / (**4_{OSi}** + CF₃H) measured *in situ* by ¹⁹F NMR after manual assembly in an NMR tube; selectivity is independent of X^- . ^btimes indicated are for >97% conversion of **2**, at 300 K. ^c85% conversion. ^d[K(L)]⁺ = K(crypt-222)⁺; generated *in situ* from KOPh + crypt-222.

Reactions where $M^+ = K^+, Cs^+$ proceeded rapidly to completion, with higher selectivity for **3_{OTMS}** / **4_{OTMS}** compared to Bu₄N⁺. Reactions stalled when the cation was Li⁺ or Na⁺.⁵³ For the K⁺-mediated system, the rate was strongly attenuated by addition of 18-crown-6, or crypt-222, with the latter causing turnover to become slower and less selective (**3_{OTMS}** / **4_{OTMS}**) than reactions initiated by TBAT (counter-cation Bu₄N⁺). The identity of M^+ was also found to affect the degree of charge-development (p ranging from 1.8 to 3.0) in the ketone ($R = Me$, Scheme 5) at the product-determining transition state for CF₃ transfer. Benzaldehydes ($R = H$) behaved analogously.

Scheme 5. Effect of initiator M^+ on reaction constant (ρ).



^ai) 4-Z-C₆H₄COR (0.2 M), **2/13** (0.2 M), **1a** (0.04 M), PhF (0.4 M), MX (0.00015 M; 0.038 mol%). Z = Ph, OMe, CF₃, Me, Br. Hammett rho values calculated from product ratios, see SI.

5. Effect of Silyl Reagent on Rate and Selectivity. To further probe the CF₃ transfer process, we compared TMSCF₃ (**1a**) with TESCF₃ (**1b**) and TIPSCF₃ (**1c**), Table 1. The effects of changing the reagent were counter-intuitive and initially misleading regarding the mechanism of CF₃ transfer, *vide infra*. Reactions employing TESCF₃ (**1b**) gave lower selectivity (**3**_{OTES} / **4**_{OTES} ≈ 1.5 / 1) and proceeded very rapidly, even at low TBAT concentrations (150 μM, 0.0375 mol%; below this, reactions failed to initiate). In contrast, reactions employing TIPSCF₃ (**1c**) proceeded very slowly, requiring high initiator concentrations to proceed efficiently (> 1.5 mM, 0.375 mol%) and gave even lower selectivity (**3**_{OTIPS} / **4**_{OTIPS} ≈ 1 / 1).

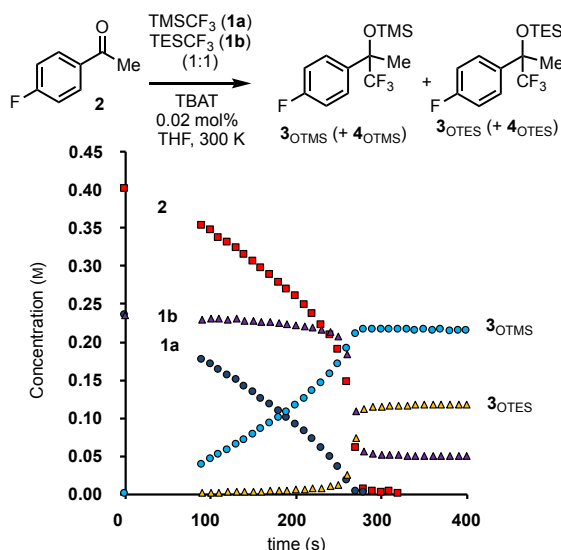


Figure 1. Competition between TMSCF₃ (**1a**) / TESCF₃ (**1b**); see text for full discussion. Reaction conditions: **2** (0.4 M), **1a** (0.24 M), **1b** (0.24 M), PhF (internal standard, 0.4 M), TBAT (75 μM, 0.019 mol%); ¹⁹F NMR analysis, manual assembly.

Further insight was afforded by reaction of a 50/50 mixture of TMSCF₃ (**1a**) and TESCF₃ (**1b**), initiated by TBAT (75 μM, 0.019 mol%), Figure 1. The first 4 minutes of reaction is dominated by turnover of TMSCF₃ (**1a**) to generate **3**_{OTMS}/**4**_{OTMS} and upon near-complete consumption of **1a**, turnover accelerates substantially as the TESCF₃ (**1b**) is engaged to generate **3**_{OTES}/**4**_{OTES}. The data indicate that the less-hindered reagent (**1a**) monopolizes the anion, but undergoes slower turnover.

Under conditions where anion-induced reactions of TMS (**1a**), TES (**1b**) and TIPS (**1c**) with **2** could be conducted slowly

enough to be accurately monitored *in situ* by ¹⁹F NMR, the ratios of enol / addition product (**4**_{OSi} / **3**_{OSi}) were all *constant* throughout the reaction evolution, Figure 2. A further distinction originated from the impact of the addition of crypt-222 to KOPh-initiated reactions. As noted above, for TMSCF₃ (**1a**) the incarceration of the K⁺ in the crypt-222 ligand substantially attenuates the rate and selectivity. In stark contrast, for TIPSCF₃ (**1c**), turnover is substantially *accelerated* by addition of crypt-222 to inhibit K⁺ / anion pairing.

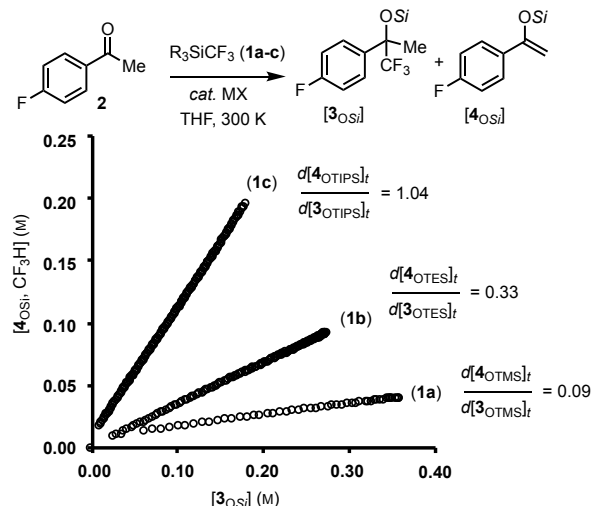


Figure 2. Constant-ratio of $[4_{\text{OSi}}]_t / [3_{\text{OSi}}]_t$. Conditions: **2** (0.4 M), **1a-c** (0.48 M), PhF (internal standard, 0.4 M), MX (TBAT 150 μM for **1a**; KOPh 0.15 mM for **1b**, TBAT 1.5 mM for **1c**).

Reactions with labelled ketone (aryl-*d*₄-**2**; CD₃-**2**; ¹³CO-**2**) were also instructive, Table 2. Reaction of TMSCF₃ (**1a**) with **2** initiated by TBAT (0.15 mM) proceeds with a very low ¹³C kinetic isotope effect (KIE), determined by competition with aryl-*d*₄-**2**, after normalizing for the effect of aryl deuteration.

Table 2. KIEs and ²H-exchange in the reaction of ketone **2**^{a,b}

CL ₃ (reagent)	1a-c	2 , 3 _{OSi}	k_H / k_D
CH ₃ (¹³ C=O)	TMSCF ₃ (1a)	CH ₃	— ($k_{12C}/k_{13C} = 1.008$) ^c
CH ₃ (C ₆ D ₄)	TMSCF ₃ (1a)	CH ₃	1.038 ^d
CD ₃	TMSCF ₃ (1a)	CD ₃	6.4 (3 / 4 _{OTMS} = 72 / 1)
CD ₃ + CH ₃	TMSCF ₃ (1a)	CD ₃ / CH ₃ only	6.1 (rate: CF ₃ H/CF ₃ D)
CD ₃	TESCF ₃ (1b)	CD ₃	3.1 (3 / 4 _{OTES} = 4.3 / 1)
CD ₃ + CH ₃	TESCF ₃ (1b)	partial CD _{3-r} /H _r	— (3 / 4 _{OTES} = 2.3 / 1)
CD ₃	TIPSCF ₃ (1c)	CD ₃	1.1 (3 / 4 _{OTIPS} = 1.1 / 1)
CD ₃ + CH ₃	TIPSCF ₃ (1c)	full CD _{3-r} /H _r	1.0 (rate: CF ₃ H/CF ₃ D)

^aKetone (**2** / ²H₃-**2**; 0.40 M), **1a-c** (0.48 M), THF, 300 K. TBAT (0.04 mol%, 0.15 mM), ^bSelectivity **3**_{OSi}/(**4**_{OSi} + CF₃H/D) and exchange measured *in situ* by ¹⁹F NMR analysis. ^cKIE determined by competition with aryl-*d*₄-**2**. ^d²HKIE induced by aryl-deuteration, determined by competition with unlabelled **2**.

In contrast, a substantial primary ^2H KIE, determined from $[\text{CF}_3\text{D}]$ versus $[d_3\text{-3Os}]$, as in Figure 2, increases the addition / enol selectivity ($k_{\text{H}}/k_{\text{D}} = 6.4$). Reactions of mixtures of **2** and $d_3\text{-2}$ proceeded with no detectable scrambling of D/H between **2** / $\text{D}_3\text{-2}$ during turnover, provided that $[\mathbf{1a}]_0 > [\mathbf{2}]_0$, and again proceeded with a high KIE ($k_{\text{H}}/k_{\text{D}} = 6.1$). With TESCF_3 (**1b**) a moderate KIE ($k_{\text{H}}/k_{\text{D}} = 3.1$) was observed, with a trace of D/H exchange between **2** and $\text{D}_3\text{-2}$ on co-reaction, and thus into products $d_n\text{-3}/d_n\text{-4}$. With TIPSCF_3 (**1a**) there was no significant KIE and a statistical mixture of isotopologues of $d_n\text{-2}/3$ ($n = 0\text{--}3$) was evident immediately after initiation of the reaction.⁵⁴

6. Variable-Ratio Stopped-Flow NMR and IR. Detailed exploration of the kinetics of the trifluoromethylation by **1a** required techniques for rapid acquisition of kinetic data (some systems had formal turnover frequencies well in excess of $5,000\text{ s}^{-1}$, *vide infra*) in a time- and material-efficient manner. Stopped-flow techniques are ideal for rapid and reproducible initiation and analysis of these reactions. However, the classic fixed-ratio dual input mode of operation ($A + B$; Figure 3a) requires separate solutions to be prepared for every variation in conditions. For a three-component process such as R_3SiCF_3 (**1**) + ketone **2** + initiator (M^+X^-), a very large number of stock solutions are required to study reactions with different concentrations of reactants and initiator.

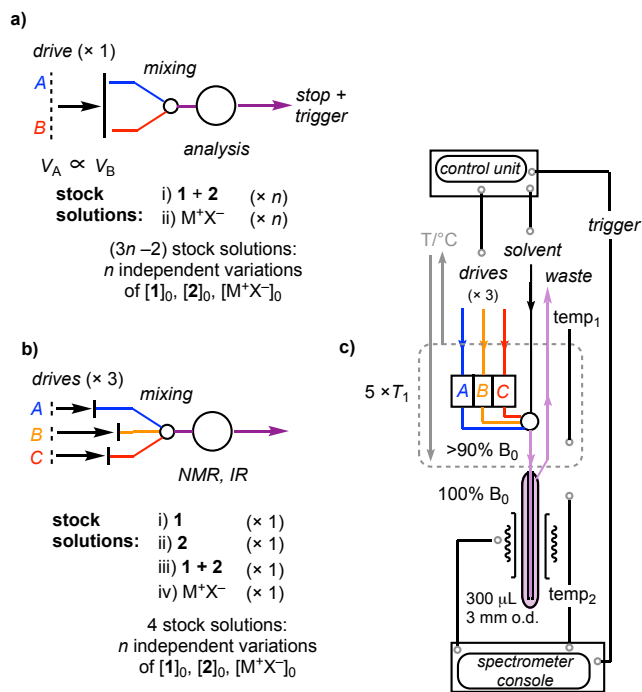


Figure 3. Schematic representations of: (a) classic fixed-ratio dual input stopped-flow; (b) a variable-ratio triple-input design; (c) variable-ratio stopped-flow NMR with thermostatic pre-magnetization of reactants (A,B,C), for $>5 \times T_1$ at $>90\%$ (B_0).

To address this issue, we constructed a stopped-flow system, in which the delivered volumes of *three* solutions (A, B, C) are independently variable,⁵⁵ using a computer-controlled triple stepper-motor system, Figure 3b. This set-up allowed systematic analysis of the kinetics across a wide range of initial

conditions, using just four stock solutions: mixing {i + iii + iv} varies $[\mathbf{2}]_0$; mixing {ii + iii + iv} varies $[\mathbf{1}]_0$; and mixing {iii + iv + THF} varies $[\text{M}^+\text{X}^-]_0$, whilst keeping the other species constant; see SI for full details. The new system was implemented in two modes: IR and NMR.⁵⁶ The former simply required adaptation of our recently developed thermostatted ATR-FTIR stopped-flow cell,⁵⁷ replacing the dual mixing stage with a triple mixer and a gated reaction volume. The analogous set-up for variable ratio stopped-flow NMR^{56a} required bespoke construction. The principles for continuous-flow NMR recently reported by Foley *et al.*⁵⁸ were employed for the basic design, such that the reaction vessel and associated components can be installed simply by insertion of the device into the sample transit of a standard unmodified NMR spectrometer. Nuclei pre-magnetization is facilitated in three independent reservoirs (A, B, C) located as close as possible to the magnetic field center, Figure 3c. The reservoirs connect at a tripodal-geometry mixer that discharges via a 0.5 mm i.d. glass capillary into a 3 mm external diameter 300 μL glass NMR flow-cell. The tube terminates at the base of the cell, with the waste outlet at the top. A fourth input to the mixer allows the system to be flushed with solvent between runs. Thermostating is achieved by passage of a heat-transfer medium (aq. ethylene glycol), using an externally controlled recirculator, through an umbilical containing all stages of the stopped-flow circuit, except for the glass flow-cell which is located within the spectrometer-thermostatted probe head; precalibration ensures $\text{temp}_1 = \text{temp}_2$. During a typical stopped-flow 'shot', a total of 600 μL is delivered through the flow-cell at a rate of $1\text{--}2\text{ mL s}^{-1}$, fully displacing the previous contents and replacing it with 300 μL of freshly-assembled reaction mixture; charging requires 70–130 msec (measured independently by UV-Vis), with high quality NMR spectra (N_2 -cryoprobe) achievable immediately thereafter. Control of the timing of the NMR pulse sequence is achieved by a trigger-signal, sent to the spectrometer console from the computer-controlled triple stepper-motor system, immediately after the 300 μL flow-cell has been freshly charged.

7. Kinetics of Trifluoromethylation by TMSCF_3 (1a**) and TIPSCF_3 (**1c**)** The kinetics of reactions initiated by M^+X^- , where $\text{M}^+ = \text{Bu}_4\text{N}^+$, K^+ , and Cs^+ , were studied in detail by SF-IR and SF-NMR across a wide range of concentrations of **1a**, **2** and $[\text{M}^+\text{X}^-]_0$. For FTIR, the decay in the IR C-F stretching mode (1056 cm^{-1}) of the TMSCF_3 (**1a**) and the growth in C-F stretching mode (1165 cm^{-1}) of **3OTMS** were collected at scan-rates of 14 or 28 s^{-1} with a resolution of 2 or 8 cm^{-1} , respectively. ^{19}F NMR analysis allowed detailed analysis of the reaction components, but was naturally more-limited in terms of temporal resolution. For faster reactions, a technique involving the interleaving of a series of spectra from a sequence of stopped-flow NMR 'shots' was employed, affording a higher virtual temporal resolution.

A key component in analysis of the kinetics was the dependence of the temporal-concentration evolution of the product (**3OTMS**) on the concentration *ratio* of ketone **2** and TMSCF_3 (**1a**). Systematic studies of initial rates using TBAT led to an

empirical rate equation for turnover frequency (TOF) in which the initiator ($\text{Bu}_4\text{N}^+\text{X}^-$) and ketone **2** are first order, and the TMSCF_3 (**1a**) reagent approximately *inverse* first order (equation 1).⁵⁹ Control experiments in which the reactions were run in the presence of exogenous product (**3**_{OTMS}) confirmed that it does not act as an inhibitor.

$$\text{TOF} \approx \frac{k_{\text{rxn}}[\text{M}^+\text{X}^-]_0[\text{2}]_t(1 - x_{\text{EI}})}{1 + K_{\text{I1}}[\text{1}]_t} \approx k_{\text{obs}} \frac{[\text{2}]}{[\text{1}]} \quad (\text{eqn. 1})$$

$$\text{TOF} \approx \frac{k_{\text{rxn}}[\text{M}^+\text{X}^-]_0[\text{1}]_t(1 - x_{\text{EI}})}{1 + K_{\text{I2}}[\text{2}]_t} \approx k_{\text{obs}} \frac{[\text{1}]}{[\text{2}]} \quad (\text{eqn. 2})$$

The inhibitory effect of the TMSCF_3 reagent **1a** (K_{I1} ; equation 1) results in very distinctive temporal concentration profiles for the reaction, simulations of which are presented later. For example, when the initial ratio of reactants is equal ($[\text{2}]_0 = [\text{1a}]_0$) their ratio remains constant ($[\text{2}]/[\text{1a}]_t = 1$) throughout the reaction. What arises is an apparent *pseudo* zero-order consumption of the reactants ($\text{TOF} = k_{\text{obs}}$) for the majority of the reaction evolution. Conversely, when there is an excess of ketone **2** over **1a** the rate of turnover rises as a function of conversion, becoming very rapid in the final phases of reaction where $[\text{2}]/[\text{1a}]_t \gg 1$.

Systematic studies using M^+X^- ($\text{M}^+ = \text{Li}^+, \text{Na}^+, \text{K}^+, \text{Cs}^+$) which induce very rapid turnover, proved more challenging. Reactions where $\text{M}^+ = \text{Li}^+, \text{Na}^+$, stalled before completion and were not reproducible. KOPh and CsOPh initiated at very low concentrations, without an evident induction period, proceeded to completion, and provided reproducible kinetics. Study of the *initial* rates suggested higher-order dependencies on TMSCF_3 ($[\text{1a}]_0$, again inverse) and on $[\text{M}^+\text{X}^-]_0$, with the ketone **2** remaining first-order. However, the reactions *evolve* with near-identical behaviour to those initiated by TBAT (equation 1).⁵⁹ The dichotomy is indicative of the presence of exogenous inhibitor(s) in low concentration in the TMSCF_3 (**1a**) reagent, that are not consumed during reaction. Increasing the initial concentration of the reagent ($[\text{1a}]_0$), or decreasing the initiator concentration ($[\text{M}^+\text{X}^-]_0$), both cause a greater mol-fraction of exogeneous inhibition (x_{EI}), equation 1.^{59,60} Addition of $[\text{K}][(\text{C}_6\text{F}_5)_4\text{B}]^-$, to provide an additional soluble K^+ source with a non-nucleophilic counter-anion, had no impact on the kinetics of reactions initiated by KOPh , indicative that the rate is dependent on the initiating anion *concentration* and the counter-cation *identity* (but not its concentration).⁶¹ Addition of potassium-binding ligands attenuated the rates substantially, and with crypt-222, the system underwent turnover slower than with Bu_4N^+ (a 3 orders of magnitude rate-reduction compared to free K^+).

The kinetics of trifluoromethylation of 4-fluorobenzaldehyde (**13**) by **1a** were also explored using TBAT as initiator. The aldehyde undergoes significantly faster trifluoromethylation than ketone **2** ($k_{\text{ald}} / k_{\text{ket}} \approx 80$, at 21 °C) requiring lower initial TBAT concentrations, and causing the traces of exogeneous inhibitor(s) in **1a** to complicate the kinetics.⁵⁹ Competing ketone **2** with aldehyde **13** (9 / 1 ratio) using stopped-flow ^{19}F NMR to analyze the transient substrate ratio (**2**/**13**) during the first 5-30 seconds of reaction indicated that the relative

rate of trifluoromethylation is independent of $[\text{TBAT}]_0$ (96–384 μM) and **1a** (0.08 to 0.48 M). Overall, the data is indicative that aldehyde **13** follows the same general kinetics as ketone **2**, i.e. equation 1.^{59,60} The rate of trifluoromethylation of ketone **2** using TIPSCF_3 (**1c**) was much slower than with **1a**. Again, the kinetics were impacted by exogenous inhibitor(s) in the reagent ($[\text{1c}]_0$), the effect of which (x_{EI}) varied from batch to batch of **1c**, see SI. Using TBAT as initiator, the reactions evolve with a first order dependency on the initiator, and on the reagent ($[\text{1c}]_t$), with inhibition by the ketone (K_{I2} ; equation 2). In other words, the kinetic dependencies are the opposite to that found for **1a** (compare equations 1 and 2) with reactions accelerating when there is an excess of **1c** over **2**. Reactions of **2** with **1c** initiated by KOPh were slower than those initiated by TBAT, and were accelerated on addition of crypt-222; the opposite phenomena to those observed with **1a**.

8. Stopped-Flow ^{19}F NMR Analysis of Siliconate and Alkoxide Intermediates, Exchange Dynamics with TMSCF_3 , and Initiator Regeneration. By use of 4-F-benzophenone (**12**; $\delta_{\text{F}} -107.0$ ppm), which reacts slower than **2**, and reducing the reaction temperature to 275 K, the temporal speciation of the initiator-derived species (10 mol% TBAT) was monitored using stopped-flow ^{19}F NMR, Figure 4. The known but unstable hypervalent bis- CF_3 -siliconate (**D**; $\delta_{\text{F}} -63.3$ ppm)³²⁻³⁴ is generated instantly. Integration against fluorobenzene (internal standard, $\delta_{\text{F}} -113.2$ ppm) shows **D** to be present at 10 mol%, and thus the predominant anion-speciation. A key feature in the time-series is the dynamic line-broadening in **D** that is constant throughout the reaction, but develops in the substoichiometric TMSCF_3 (**1a**) reagent ($\delta_{\text{F}} -66.6$ ppm) as its concentration is depleted by the overall reaction with ketone **12**. In parallel with this is a marked acceleration in product generation (**14**_{OTMS}, $\delta_{\text{F}} -72.4$ and -113.7 ppm), consistent with equation 1. After 6 seconds, the TMSCF_3 (**1a**) is fully consumed and TBAT ($\delta_{\text{F}} -97.4$ ppm) is regenerated from $\text{Ph}_3\text{SiF} / \text{Me}_3\text{SiF}$. The dynamic-line broadening in **D** / **1a** can be satisfactorily simulated using a 3-spin exchange process in which **D** is in rapid dissociative equilibrium ($k_{\text{exch}} \sim 180 \text{ s}^{-1}$; $\Delta G^\ddagger \sim 13 \text{ kcal mol}^{-1}$) with **1a** and a low concentration of (unobserved) $[\text{Bu}_4\text{N}][\text{CF}_3]$ (**E**).⁶² At 300 K, the line-broadening is very extensive and **D** short-lived.

Analogous experiments using TIPSCF_3 (**1c**) gave a very different outcome. Reactions conducted with **1c** at 275 K were slow enough to be followed using ketone **2** ($\delta_{\text{F}} -106.7$ ppm), Figure 5. The ^{19}F NMR signal for **1c** remains sharp until **2** has been fully consumed. In contrast to reactions with **1a** (Figure 4) the alkoxide (**3**_{O-}; $\delta_{\text{F}} -118.4$) is present in significant concentration and exhibits dynamic-line broadening (see inset to Figure 5). The signal for ketone **2**, also exhibits dynamic-line broadening, immediately after addition of the TBAT. On complete consumption of **2** (~ 120 s) the signals for remaining **1c** and CF_3H are broadened, presumably due to indirect exchange involving CF_3^- . After a further 300 s, **1c** is fully consumed and the CF_3H doublet becomes sharp again.

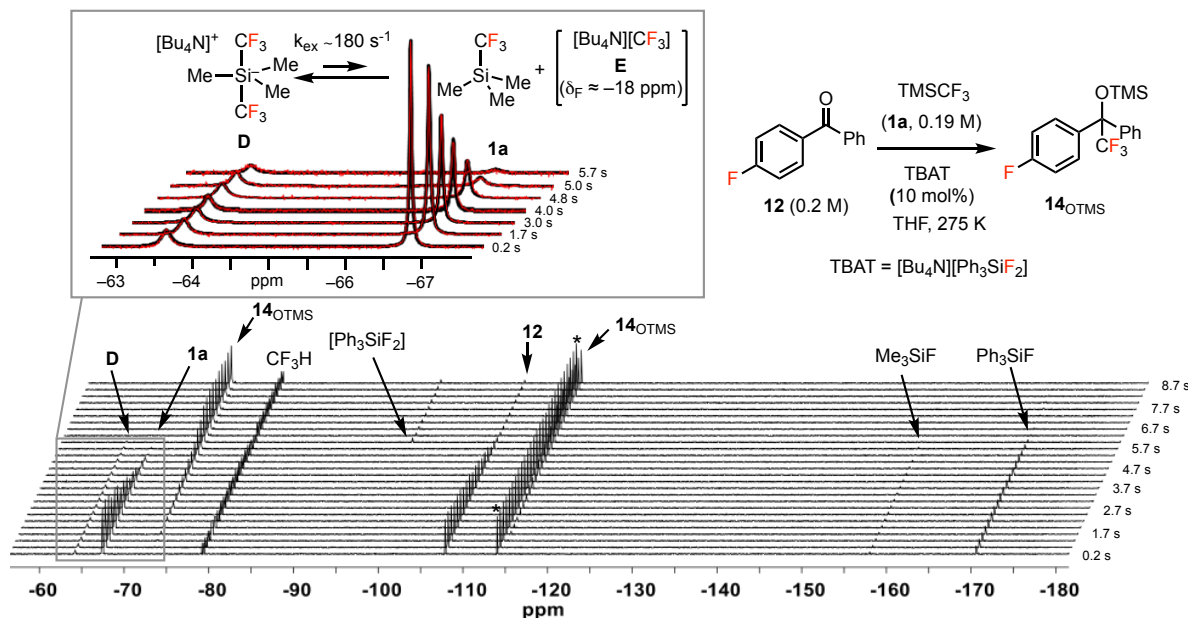


Figure 4. Selected spectra from stopped-flow ^{19}F NMR analysis of the reaction of 4-F-benzophenone **12** (0.20 M) with **1a** (0.19 M) in THF at 275 K after initiation by 10 mol% TBAT. Inset: overlay of selected simulations⁶² (black) of dynamic line-shape for **D** / **1a** with experimental spectra (red); **E** ($\delta_{\text{F}} \sim -18$ ppm) is undetected. (*) $\text{C}_6\text{H}_5\text{F}$ internal standard.

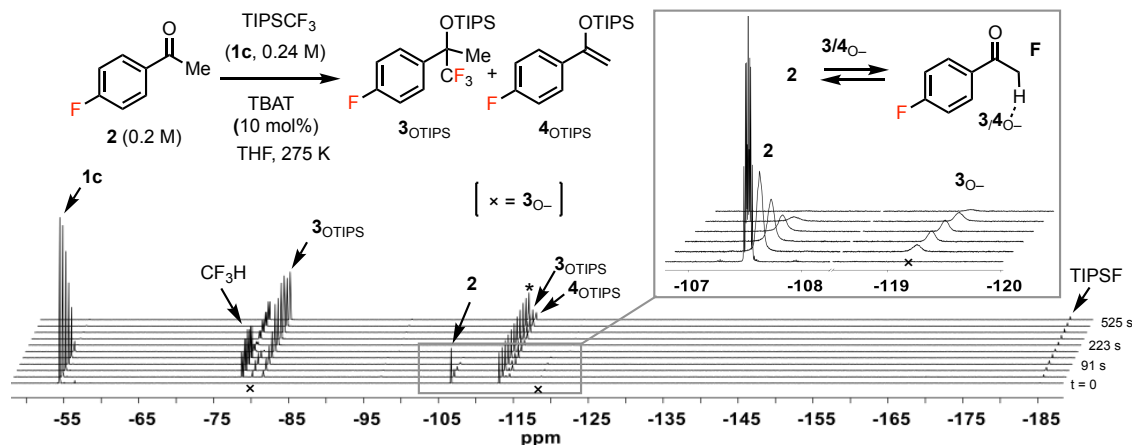


Figure 5. Selected spectra from *in situ* ^{19}F NMR analysis (manual assembly) of the reaction of 4-F-acetophenone **2** (0.20 M) with **1c** (0.2 M) in THF at 275 K after initiation by 10 mol% TBAT; ($t = 0$, no TBAT). Inset: line-broadening in ketone **2** and alkoxide **3O-**. (*) $\text{C}_6\text{H}_5\text{F}$ internal standard. (x) = **3O-**. Free **4O-** not located, possibly due to degenerate exchange with **2**. Ph_3SiF is not observed.

9. General Mechanism for Anion-initiated CF_3 Transfer from R_3SiCF_3 to Ketones and Aldehydes. The data outlined in Sections 2 to 8 above (see SI for full details) indicate that the M^+X^- initiated trifluoromethylation of ketone **2** by TMSCF_3 (**1a**) involves an electrophile-nucleophile reaction, in which the CF_3 transfer is accompanied by M^+ . Enolsilane **4OTMS** is also generated ($\leq 2\%$ when $\text{M}^+ = \text{K}^+$ and 7% when $\text{M}^+ = \text{Bu}_4\text{N}^+$) with co-product CF_3H ($k_{\text{H}}/k_{\text{D}} = 6.1$). Using TIPSCF_3 (**1c**), approximately 50% of the product is **4OTIPS** and $k_{\text{H}}/k_{\text{D}} = 1.0$. Contrasting kinetic behaviour is observed for **1a** (equation 1) versus **1c** (equation 2), with the roles of reactant for turnover, and inhibitor reversed between the two systems. These disparate sets of observations can easily be misinterpreted as turnover for **1a** versus **1c** arising from different pathways, e.g. silicate versus carbanion. However, analysis of the kinetics, KIEs, and DFT calculations of a wide range of potential intermediates (see SI), eventually leads to the conclusion that the two reagents elicit contrasting kinetics,

selectivity (**3osi**/**4osi**) and KIEs, by biasing one of two extremes in a single overarching mechanism. Calculations employed the M06L/6-31+G* level of theory, which was selected from a range of other functionals and larger basis sets that were considered,⁶³ see SI, as it provided the best quantitative agreement with experiment. All calculations were performed in Gaussian09,⁶⁴ with THF solvation incorporated via a PCM single point at the same level of theory, and with $T = 298$ K and pressure at 24.45 atm to achieve a 1 M standard state.⁶⁵ Kinetic isotope effects were computed using the Kinisot program.⁶⁶ Some of the TES and TIPS bearing structures required the 'loose' settings during the geometry optimization, presumably because of the flat potential energy surface associated with the long Si- CF_3 bonds.

The calculations permitted several possible structure types (such as hexacoordinate silicon dianions) to be excluded from consideration, and also revealed pronounced differences between intermediates based on TMS, TES, and TIPS,

where the increasing steric bulk substantially destabilizes the pentacoordinate anions, Figure 6. Extensive calculations were conducted to test for direct nucleophilic reactivity of the pentacoordinate anions **B** and **D**. All calculations revealed that *direct* transfer of CF₃ from the silicon centre to an electrophile requires concomitant inversion of the CF₃, with a prohibitively large barrier (>100 kcal mol⁻¹; in line with the barrier for inversion of the free CF₃ anion).⁶⁷ The pentacoordinate silicate anions thus act as reservoirs, not active nucleophiles, liberating free (non-silicon coordinated) CF₃⁻ via dissociation. The transition state for addition of the CF₃⁻ anion(oid) to the ketone formally involves movement between a non-classical hydrogen bonded complex and the addition product, a process that occurs with low calculated barrier (7.5 kcal mol⁻¹) and well-represents the process that occurs once the two species are in contact. The calculations support the known preference for deprotonation (*k*_{CH}) in the gas phase,⁶⁸ and for addition (*k*_{CO}) once solvation is introduced,

as observed experimentally for TMSCF₃ (**1a**). The loose addition transition state leads to a negligible ¹³C KIE (carbonyl) for addition, while a large primary ²H KIE is computed for C–H deprotonation. Relative rates computed from activation free energies suggest $\rho = 2.0$ for addition to acetophenones, and a lower barrier for addition to 4-F-benzaldehyde (**13**) versus **2** ($\Delta\Delta G^\ddagger$ 2.6 kcal/mol; $k_{\text{rel}} = 81$). All of these computed values are in excellent agreement with experiment.

A general mechanism for the trifluoromethylation of ketones and aldehydes by R₃SiCF₃ reagents (**1**) in the presence of a catalytic quantity of initiator (M⁺X⁻) can thus be assembled, Figure 7. The one overarching mechanism, discussed below in the context of two extremes (Vi and Vii), rationalizes why the turnover rate (per M⁺X⁻ initiator) for a given electrophile depends on the initial concentration (but not identity) of X⁻, the identity (but not concentration) of M⁺, the identity of the reagent (**1a-c**), and the electrophile / reagent ratio (**2** / **1**).

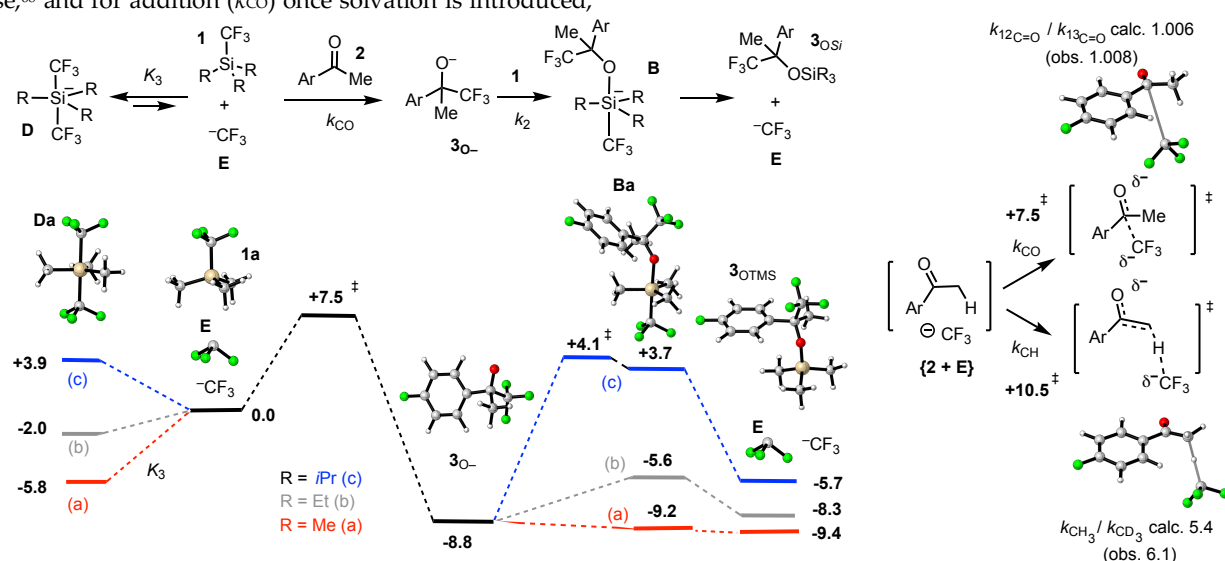


Figure 6. Selected structures and energies (M06L/6-31+G*; PCM (THF); standard state, 1 M; 298 K) of naked anions in the reaction of ketone **2** with R₃SiCF₃ **1a-c**. Energies have been normalized to [CF₃⁻ + **1** + **2**] = 0.00 kcal mol⁻¹. See text, Figure 9, and the SI for discussion of the binding modes and effects of cations. The structures and energies of other potential intermediates examined, including hexacoordinate dianions and fluoride adducts, are provided in the SI.

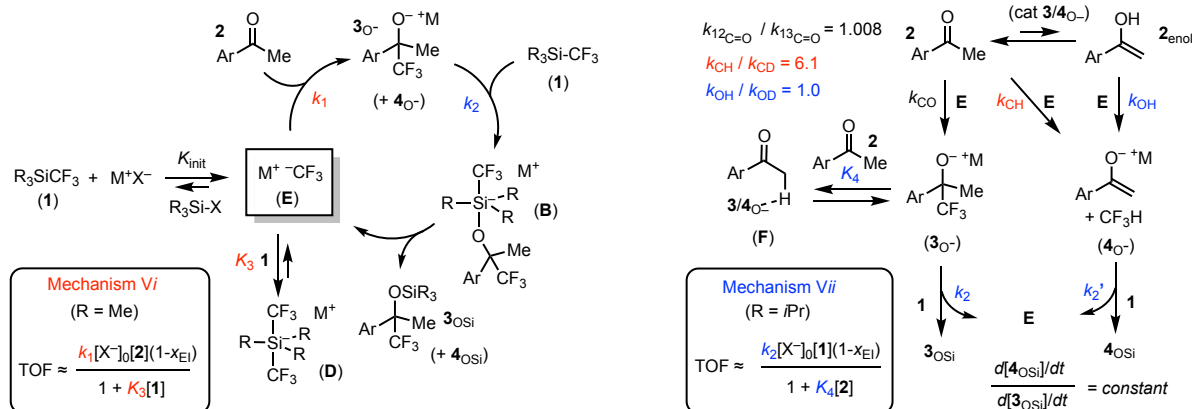


Figure 7. Mechanisms Vi and Vii: two extremes of general model V for the trifluoromethylation of ketones by R₃SiCF₃ reagents **1a-1c**, in the presence of a catalytic quantity of initiator M⁺X⁻, with acetophenone as a generic reactant. Turnover frequency (TOF) equations are simplifications of a global approximation, where $k_1 = k_{\text{CO}} + k_{\text{CH}} + k_{\text{OH}}[\mathbf{2}_{\text{enol}}]/[\mathbf{2}]$, and the mol-fraction of active anion quenched by unidentified exogenous inhibitor(s) in **1**, is χ_{EI} . Initiation (K_{init}) is not included in the rate equation. When M⁺X⁻ is 'TBAT', initiation is reversible using **1a**. For non-enolizable ketones and aldehydes, k_{CH} , k_{OH} , and $K_4 = 0$.

10. Mechanism Vi. In this regime, which describes reactions involving TMSCF_3 (**1a**), the dominant anion speciation is the bis(trifluoromethyl) siliconate (**D**),³²⁻³⁴ generated in rapid equilibrium (K_3) with CF_3^- (**E**)^{36,38} and **1a**, as observed by NMR, Figure 4. The product-determining step (k_1) involves reaction of CF_3^- (**E**) with the ketone (**2**), ($k_{\text{CO}} + k_{\text{CH}}$), and the reagent (**1a**) thus acts as a reversible inhibitor. The stronger the association of M^+ with CF_3^- (see Section 13) and with the carbonyl oxygen, the faster the turnover rate: $\text{Bu}_4\text{N}^+ < [\text{K}(\text{crypt-222})]^+ < [\text{K}(\text{18-c-6})]^+ < \text{K}^+$. The initial concentration ratios of the reactant versus the reagent dictate the temporal evolution of the reaction. When $[\text{2}]_0/[\text{1a}]_0 = 1$, *pseudo* zero order kinetics are obtained, whereas when $[\text{2}]_0/[\text{1a}]_0 \geq 1$ the rate rises throughout the reaction, becoming very fast (asymptoting to $k_3[\text{D}]$) in the final stages. The kinetics of trifluoromethylation of ketone **2** by TMSCF_3 (**1a**), can be satisfactorily simulated, Figure 8, using a truncated form of mechanism Vi that retains relationships required for TOF modulation as the temporal concentration ratio $[\text{2}]_t/[\text{1a}]_t$ evolves.

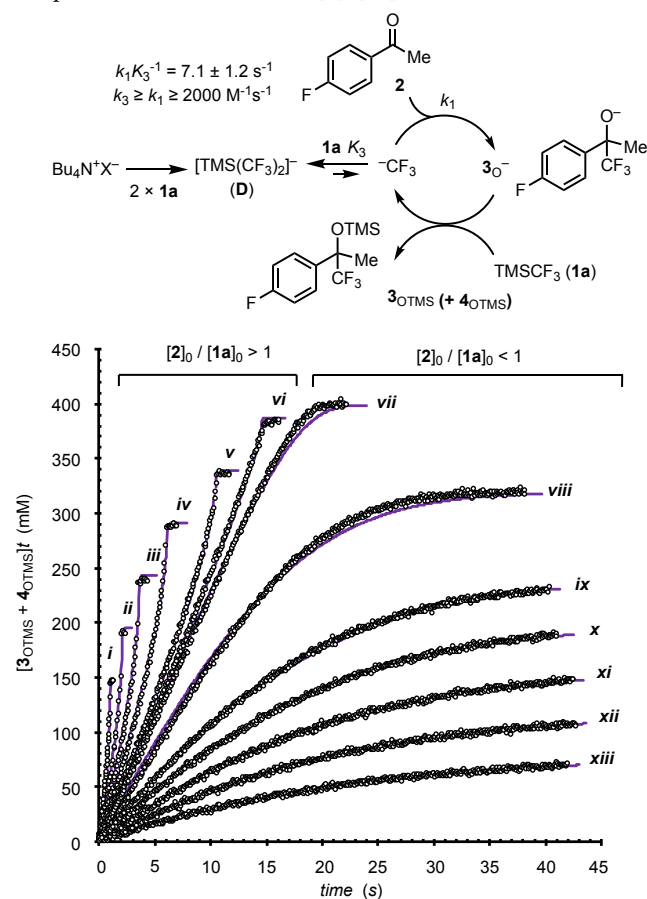


Figure 8. Simulation of experimental data (open circles, SF-IR; $[\text{3OTMS} + \text{4OTMS}]_t$) based on simplified mechanism Vi, for reaction of ketone **2** with TMSCF_3 (**1a**), initiated by 3.6 mM TBAT ($\text{Bu}_4\text{N}^+\text{X}^-$). For $[\text{2}]_0/[\text{1a}]_0 > 1$, $[\text{2}]_0 = 0.40$ M and $[\text{1a}]_0 = 144, 192, 248, 288, 336, 384$ mM (i to vi). For $[\text{2}]_0/[\text{1a}]_0 < 1$, $[\text{1a}]_0 = 0.48$ M and $[\text{2}]_0 = 400, 320, 240, 200, 160, 120, 80$ mM (vii to xiii). Induction and turnover by **1a** are set to arbitrary high values. Fitted parameters (k_1 , k_3 , k_{-3}) as indicated; $x_{\text{EI}} = 0$.⁵⁹

11. Mechanism Vii. In this regime, which describes reactions involving TIPSCF_3 (**1c**), the dominant anion speciation

is a combination of the product alkoxide (**3o-**), the enolate anion (**4o-**), and MX . Ketone (**2**) can reversibly H-bond (**F**) with oxy-anions **3/4o-**, as observed by NMR, Figure 5, leading to inhibition (**K4**).⁶⁹ When $[\text{1c}]_0/[\text{2}]_0 = 1$ *pseudo* zero order kinetics are observed; reactions in which $[\text{1c}]_0/[\text{2}]_0 > 1$ exhibit accelerating rate in the last stages of reaction. The more strongly bound M^+ to **3/4o-**, the slower the reaction with **1c**, leading to rates increasing in the series: $\text{K}^+ < [\text{K}(\text{18-c-6})]^+ < \text{Bu}_4\text{N}^+ < [\text{K}(\text{crypt-222})]^+$; i.e. the opposite order to Vi. When the non-enolizable ketone 4-F-benzophenone **12** is employed, the kinetics show clean *pseudo* first-order decay in **1c**, see SI, with no inhibition by **12** (i.e. Mechanism Vii, where $K_4 = 0$; and equation 2, where $K_{i2} = 0$).

12. Competing Enolization. Also shown in Figure 7 is the generation of the enol ether (**4osi**) and CF_3H from ketone **2**, for which the selectivity (**4osi** / **3osi**) is dependent on M^+ and the reagent (**1a-c**), Table 1. The major pathway for generation of **4OTMS** in mechanism Vi is via C–H deprotonation (k_{CH}) with an attendant large primary ^2H -KIE.^{70,71} In contrast, for mechanism Vii, the significant concentration of **3/4o-** allows keto-enol equilibrium ($pK_{\text{enol}} \sim 8$)⁷² in **2** to be approached, with attendant intermolecular scrambling of ^2H between ketone methyl groups. Deprotonation (k_{OH}) of the enol (**2enol**) is predicted (DFT) to be of very low barrier, and thus proceed with a negligible ^2H -KIE.⁷⁰ Despite their different origins (k_{CH} versus k_{OH}) mechanisms Vi, and Vii both lead to **4osi** / **3osi** ratios that are independent of the concentration of reactants (**1**, **2**) and constant throughout reaction, Figure 2.

13. Cation- CF_3^- Interactions. The interactions between the CF_3^- anion (free and Si-bound) and the counter-cations K^+ and Me_4N^+ (as a model for Bu_4N^+) were explored computationally, with multidentate CF_3^- interactions found to be favored, e.g. Figure 9; see SI for details.

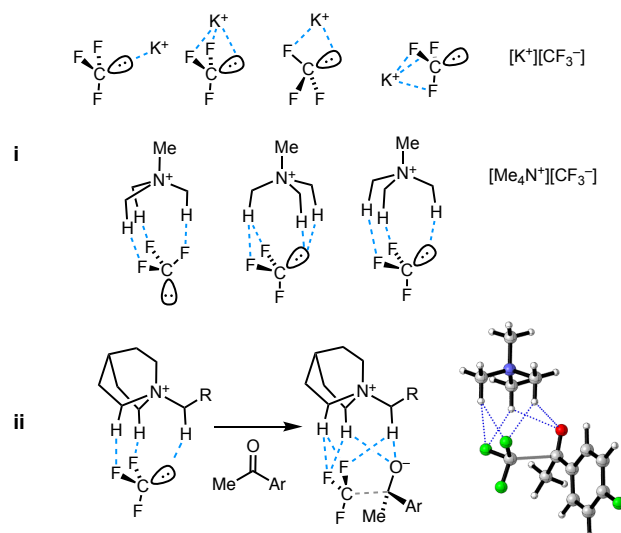


Figure 9. Cation binding to free CF_3^- anion: (i) various modes of binding of K^+ and Me_4N^+ cations; (ii) concept (schematic) for enantioselective addition beneath the quinuclidinium core of a cinchonidinium initiator. Inset: structure of TS for addition of $[\text{CF}_3^-][\text{Me}_4\text{N}^+]$ to **2**, see SI, with H-bonding interactions to developing alkoxide anion.

The *indirect* transfer of CF_3 from reagent **1a** to the ketone / aldehyde, i.e. via a silicon-free carbanion **E**, has implications

for the mode by which enantioselective catalysis can be achieved using chiral ammonium initiators, e.g. cinchonidinium salts. The CF₃-anion binding modes found computationally for Me₄N⁺ (Figure 9i) show how an ammonium cation might simultaneously interact with a CF₃[−] anion and control a developing alkoxide anion, Figure 9ii. Mechanism Vi contrasts most,^{16de,31} but not all,^{16f} prior interpretations, where mechanisms II/III (Scheme 2) involving CF₃-siliconates bearing the initiating (C) or propagating (B) anion, are proposed to play key roles in the enantioselective trifluoromethylation step.

14. Broader Mechanistic Aspects. The mechanistic features elucidated in the current study extend beyond carbonyl trifluoromethylation. A number of corollaries follow for generic anion-initiated trifluoromethylation of an electrophile (E) by **1a**, or deprotonation (R–H),⁷³ via pathways analogous to mechanism Vi, and where [E, R–H]₀ >> [M⁺X[−]]₀. Thus, the initiator (M⁺X[−]) affects the rate of reaction in a number of ways. [X[−]]₀ sets the initial concentration of the siliconate ([D]₀ = (1 − x_{EI})[X[−]]₀)⁶⁰ which, in the absence of endogenous inhibitors, is essentially constant throughout reaction. The insurmountable barrier for CF₃ inversion⁶⁷ means that, independent of the identity of the electrophile, E, or proton donor, R–H, the siliconate is unable to effect direct anionic trifluoromethyl transfer, Figure 10i. In all cases, the reaction must proceed via a dissociative pathway, Figure 10ii, in which M⁺ plays a key role: the stronger the association of M⁺ with CF₃ the more favourable *k*_{−3}. In contrast, efficient regeneration of the siliconate (*k*₂, Figure 7) is favored by weaker interactions between M⁺ and the anionic co-product from trifluoromethyl transfer (CF₃–E[−]; R[−]; or products thereof). When the anion is unable to react with **1a**, stoichiometric initiation by [M⁺X[−]] is required.^{14–26}

15. Exogenous Inhibition. Trifluoromethylations initiated by low concentrations of (M⁺X[−]) are highly sensitive to traces of exogenous inhibitor(s). Species that generate of an anion (LG[−]) of insufficient reactivity towards **1a** to propagate, will terminate the anionic chain reaction, Figure 10iii. In a series of control experiments, additives of the form Z–LG, (Z = H, R₃Si, LG = Cl, Br) were found to function as powerful inhibitors for the anion-initiated reaction of **2** with **1a**. For example, the trifluoromethylation of **2** (0.4 M) initiated by 150 μM TBAT ceases immediately on addition of 150 μM TMSCl, see SI. Slower-onset irreversible inhibition is effected by the more hindered TIPSCl, which is also inhibits the reaction of TIPSCF₃ (**1c**). Competing consumption of **1a** is effected by other species in low-concentrations, including CCl₄ (Cl-transfer),⁷⁴ Cl₃CH (deprotonation / Cl-transfer),^{73b} and TMS–OH (deprotonation), but without significant chain termination. There was no detectable inhibition by DCE, CH₂Cl₂,^{73b} TMS–O–TMS, Ph₃SiF, Me₃SiCF₂H, or MeCN.^{73c}

In our experience, a diverse range of inhibitors, and competitors (e.g. CCl₄ and CHCl₃) are present, in low concentrations and variable proportions, in commercial samples of TMSCF₃ (**1a**). This leads to substantial differences in reaction outcome, depending on the supplier. For example, comparison of the reaction of **2** (0.40 M) with five samples of freshly-distilled **1a** (0.48 M) revealed that the concentration of initiator

(TBAT, KOPh) required to effect >99% conversion of **2** ranged from 30 μM to 2.0 mM (0.0075 to 0.5 mol%); see SI.

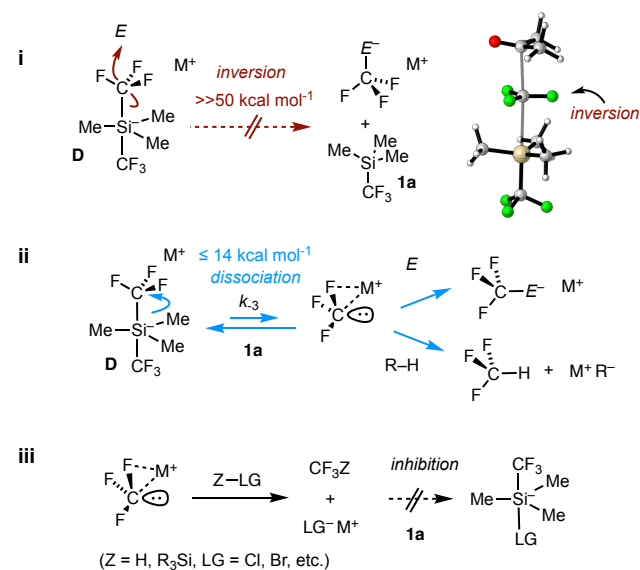


Figure 10. Generic reactivity of siliconate **D** towards electrophiles (E), carbon acids (R–H), and inhibitors (Z–LG). (i) direct CF₃ transfer from **D** is strongly disfavoured. Inset: TS for CF₃-transfer to acetone, see SI; (ii) dissociative CF₃ transfer, without CF₃ inversion; (iii) termination of the anionic chain reaction by traces of exogenous inhibitor(s), or substrates that generate an unreactive anion, LG[−].

An major difference found between reactions involving reagent **1a** versus **1b,c** is the impact of the persistent radical, TEMPO, which powerfully inhibits reactions involving **1a**, Scheme 4A. The difference in behavior towards TEMPO cannot arise from oxidation of the CF₃ anion (E), as this is a common intermediate to all three reagents (**1a–c**), and the partitioning of E between reaction with the ketone (**2**) versus TEMPO will be constant across the series, i.e. independent of the provenance of the carbanion E. Since the major difference between reagents **1a** and **1c** under the conditions of the reaction, is the dominant anion speciation (**D**; mechanism Vi, **1a**, versus alkoxides **3o**[−] / **4o**[−]; mechanism Vii, **1c**), this suggests that reaction of siliconate **D** with TEMPO is responsible for the inhibition. We were unable to identify any products *in situ*, or by quenching, arising from TEMPO under the standard reaction conditions, see SI. Whilst siliconates of type **D** are also generated from **1b** and **1c**, they are a) only present as low concentration or transient species, thus reducing their net rate of reaction with TEMPO, and b) may be more resistant to reaction with TEMPO due to their greater steric bulk.

CONCLUSIONS

The trifluoromethylation of ketones and aldehydes by TMSCF₃ (**1a**), initiated by catalytic fluoride ion, has been employed in synthesis for three decades.¹² Previous mechanistic work has focussed on stoichiometric reactions of R₃SiCF₃ (**1a,c**) with anions at low temperatures, generating unstable trifluoromethyl siliconates (**C,D**)^{32–34} and carbanion(oids) (E),^{36–38} depending on conditions. Which of these two pathways is followed in catalytic reactions at ambient temperature has been a long-standing mechanistic dichotomy.³⁰ A

variable-ratio stopped-flow NMR/IR approach (Figure 3) has been developed to facilitate time- and material-efficient analysis of a wide range of initiator (M^+X^-) and reactant concentrations. Change of reagent from $TMSCF_3$ (**1a**) to $TIPSCF_3$ (**1c**) has a profound impact on the reaction. For example, the conversion of 4-F-acetophenone (**2**, 0.4 M) to **3**_{OTMS} by equimolar **1a** in THF at ambient temperature takes < 125 msec to complete using 0.1 mol% KOPh initiator, and generates <2% of silylenol ether **4**_{OTMS}, whereas with $TIPSCF_3$ (**1c**) and 3.75 mol% KOPh, the reaction proceeds to just 60% conversion in 16 hours, and generates 50% **4**_{OTIPS}. The rates of reaction are strongly affected by traces of inhibitors present in the reagents (**1**), especially at the low concentrations of initiator (M^+X^-) employed for the fastest reacting systems, see equations 1 and 2.^{59,60} Nonetheless, whilst these render initial rate data misleading, study of the full reaction time-course, e.g. Figure 8, provides a coherent kinetics analysis.

A unified mechanism (V) for the reaction of R_3SiCF_3 reagents (**1a-c**) with ketones and aldehydes under conditions of catalytic anionic initiator (M^+X^-) is presented in Figure 7. The work confirms that the carbanion³⁶⁻³⁸ mechanism prevails under conditions of application (Scheme 1). Mechanism V allows a number of initially confusing observations to be rationalized. The main difference between use of $TMSCF_3$ (**1a**) versus $TIPSCF_3$ (**1c**) reagents is an inversion in the major anion speciation in the overall anionic chain reaction. This inversion leads to opposing influences of electrophile and silicon reagent (mechanisms *Vi* and *Vii*), and to keto-enol equilibration (**2** / **2**_{enol}) with **1c** (*Vii*). When TBAT is used as initiator,⁷⁵ $TESCF_3$ (**1b**) effects the most rapid trifluoromethylation in the series **1a-c**. The increased steric bulk in **1b** reduces reagent inhibition (K_3) relative to **1a**, without the substantial kinetic penalty in k_2 experienced by **1c**. These factors shift the reaction with **1b** closer to an 'ideal' catalytic cycle in which the intermediates are all connected by low TS barriers, with reduced off-cycle speciation. A consequence of adding

$TMSCF_3$ (**1a**) to $TESCF_3$ (**1b**) is therefore to strongly inhibit turnover of **1b** until all of **1a** has been consumed, Figure 1.

The overarching mechanism (V, Figure 7) for anion-initiated reactions of R_3SiCF_3 (**1**) with ketones and aldehydes should prove of utility in their application in synthesis. For example, in the context of the design and analysis of enantioselective trifluoromethylation processes,^{16,29,30a,31} mechanism V shows that control must be achieved by the $CF_3^- / [M]^+$ ion pair, Figure 9ii, and not by a siliconate intermediate. Moreover, the key mechanistic features of the anion-initiated reactions of **1** with carbonyl compounds (Figure 7) translate to reactions of **1** with other electrophiles (*E*),²⁹⁻³¹ and proton donors ($R-H$ to generate R^-),⁷³ Figure 10. Thus, all processes in which siliconate **D** or analogous species, formally acts as a nucleophilic or basic source of CF_3^- , must proceed via a dissociative pathway (Figure 10ii). Siliconate **D** is inherently unstable, and decomposes at ambient temperature to generate, inter alia, complex perfluorocarbanions.^{34,38a} The rate of anionic chain transfer, as dictated by the reactivity of the electrophile (*E*)²⁹⁻³¹ or carbon acid ($R-H$)⁷³ towards CF_3^- , as well as the presence of species able to attenuate decomposition (e.g. via CF_2 -capture, **4**_{OSi} → **10**, Scheme 3), controls the formal lifetime of **D**, and in turn the minimum loading of initiator (M^+X^-) that will be required to achieve complete conversion of substrate. Moreover, traces of exogenous inhibitor(s), (e.g. Z-LG, Figure 10iii) ubiquitous in R_3SiCF_3 reagents (**1**), act to reduce the net active anion in the chain reaction, again increasing the requisite loading of initiator (M^+X^-). Compounds employed in synthetic routes to reagents **1a-c**, e.g. $TMSCl$ and $TIPSCl$,¹¹ function as powerful inhibitors. However, the identity and effect of the inhibitors in reagents **1a-c** vary substantially from batch to batch, and between commercial suppliers (see SI). Electrophiles or carbon acids ($R-H$) that react with CF_3^- to ultimately generate an anion of inherently low reactivity towards **1**, require stoichiometric initiator to proceed to completion.¹⁴⁻²⁶

ASSOCIATED CONTENT

Supporting Information: Additional discussion, experimental procedures, further kinetic data and analysis, characterization data and NMR spectra. This material is available free of charge via the Internet at <http://pubs.acs.org>.

AUTHOR INFORMATION

Corresponding Author

Guy.lloyd-jones@ed.ac.uk

Author contributions

† C.P.J. and T.H.W. contributed equally to this work.

Notes

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- (60) In equations 1 and 2, and elsewhere, $(1-x_{\text{EI}})$ represents the mol-fraction of active anion relative to total anion $[M^+X^-]_0$. Based on a 1:1 inhibition mode, $x_{\text{EI}} = (K_{\text{EI}}[I]/(1+K_{\text{EI}}[I]))$ with $[I]_0 = x_i[1]_0$, where x_i = mole fraction inhibitor in reagent **1**. Experimental data (see SI) suggest K_{EI} is substantially greater with **1c**.
- (61) This contrasts the borazine systems recent developed by Szymczak (see ref 10) where addition of KBARF profoundly accelerates CF₃-transfer rates.
- (62) Simulations were conducted using the three-spin parametrization in WNDNMR, with the rates and frequencies as indicated in Figure 4 ($k_{\text{ex}} = 2k_{\text{ex}}$; where k_{ex} is the apparent ¹⁹F nuclei exchange rate through reassociation). The fraction of CF₃ present as **E** was arbitrarily set to 0.1%, with the remaining 99.9% partitioned between **D** and **1a** as required to fit. The chemical shift of CF₃[M] is reported as $\delta_{\text{F}} - 18.7$ ppm, see reference 38, and -17.2 ppm, see reference 36, depending on the identity of [M].
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GRAPHICAL ABSTRACT

